

Science 9: Unit 1: Chemistry
Test Review

Materials taken from the following sections of the textbook:

Section 1.1- safety symbols, part 2 Setting safety rules questions
Section 1.2- Properties of Matter- physical properties, chemical properties and alloy
Section 1.7- Physical and Chemical Change- clues of a chemical change
Section 1.9- Corrosion- preventing corrosion
Section 1.11- Combustion- word equation, fire triangle
Section 2.1- Pure substances and mixtures – particle theory, elements/compounds, atoms/molecules.
Section 2.7- Chemical symbols/formulas
Section 3.3- Inside the Atom, charged atom (ion)
Section 3.4- Bohr/Bohr Rutherford Diagrams
Section 3.6- Isotopes/radioactivity
Section 4.4- Group of Elements

Make sure you are able to define the following words:

Matter	State	Physical properties	Chemical properties
Hardness	Malleability	Melting point	Boiling point
Crystal form	Solubility	Viscosity	Density
Combustible	Reaction with acid	Alloy	Physical change
Chemical change	Pure substances	Mixtures	Solution
Elements	Compounds	Atom	Molecule
Combining capacity	Ductile	Heterogeneous & Homogeneous mixtures	
Ion	Corrosion	Proton	Electron
Neutron	Sublimation	Element	Compound

Know:

The differences between physical and chemical properties and be able to identify each. (In Multiple choice)

WHMIS and HHPS symbols and definitions

The differences between physical and chemical changes and be able to identify each.(In Multiple Choice)

Particle Theory of matter (4 parts)

Be able to:

Count atoms in compounds

Tell the difference between chemical symbols and formulas

Draw Bohr-Rutherford Diagrams

Locate Elements on the periodic table given group and period (Ex. What element is Group5 Period 4?)

Explain “All compounds are molecules but not all molecules are compounds”

Use the periodic table to fill in the charts for Element Name, Atomic Number, Mass Number, Standard Notation, Numbers of protons, Number of electrons. Similar to assignments.

15 Matching of terms, 15 Multiple Choice, 8 points for identifying chemical and physical changes, Know diagrams from class for pure element, pure compound, mixture of elements, mixtures of element and compound, and mixture of compounds (5points). Explain why “all compounds are molecules but not all molecules are compounds” (3points), Chart of Protons/Neutrons/Electrons/Mass number/atomic number (11 points), Counting atoms (7 points), 2 Bohr diagrams (8 points)